

MICROCOPY RESOLUTION TEST CHART
NATIONAL BUREAU OF STANDARDS-1963-A

OFFICE OF NAVAL RESEARCH

Contract N00014-80-C-0472

Task No. NR 056-749

TECHNICAL REPORT No. 56

Phase-Dissipative Mechanisms for Laser-Induced Surface Desorption/Dissociation Processes

by

Xi-Yi Huang, Thomas F. George, Jian-Min Yuan and L. M. Narducci

Prepared for Publication

in

The Journal of Physical Chemistry

LE COPY

Department of Chemistry University of Rochester Rochester, New York 14627

October 1984

1 = 1 84

This document has been approved for public release and sale; its distribution is unlimited.

Unclassified

| a Entered) | |
|---|---|
| REPORT DOCUMENTATION PAGE | |
| | 3. RECIPIENT'S CATALOG NUMBER |
| 4. TITLE (and Sublice) Phase-Dissipative Mechanisms for Laser-Induced Surface Desorption/Dissociation Processes | |
| ian-Min Yuan | CONTRACT OR GRANT NUMBER(*) N00014-80-C-0472 |
| .· | 10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS NR 056-749 |
| and from Constanting Office. | 12. REPORT DATE October 1984 13. NUMBER OF PAGES 11 ffice) 15. SECURITY CLASS. (of this report) |
| mi from Controlling Office) | Unclassified 15a. DECLASSIFICATION/DOWNGRADING SCHEDULE |
| | Laser-Induced Processes |

e contraction of the contraction

This document has been approved for public release and sale; its distribution is unlimited.

17. DISTRIBUTION STATEMENT (of the obstract entered in Block 20, If different from Report)

18. SUPPLEMENTARY NOTES

Prepared for publication in the Journal of Physical Chemistry.

19. KEY WORDS (Continue on reverse side if necessary and identify by block number)

ADMOLECULE-SURFACE SYSTEMS LASER-INDUCED VIBRATIONAL EXCITATION **ENERGY-RELAXATION MECHANISMS** PHASE-RELAXATION MECHANISMS

DESORPTION DISSOCIATION QUANTUM MECHANICAL STUDY ZWANZIG PROJECTORS

ABSTRACT (Continue on reverse side if necessary and identify by block) SVibrational excitation and relaxation for a molecule adsorbed on a surface is investigated in connection with desorption/dissociation processes. The population equations for a laserdriven anharmonic oscillator involving both energy- and phase-relaxation mechanisms are derived with the Zwanzig projector technique and eigenfunctionexpansion method due to Weidlich. Phase relaxation is seen to assist in the excitation of the resonant active mode.

DD 1 JAN 73 1473

2 PROCESSES

EDITION OF 1 NOV 68 IS OBSOLETE 5/N 0102-LF-014-6601

Unclassified

SECURITY CLASSIFICATION OF THIS PAGE (Then Date &

Journal of Physical Chemistry, in press

PHASE-DISSIPATIVE MECHANISMS FOR LASER-INDUCED SURFACE DESORPTION/DISSOCIATION PROCESSES

Xi-Yi Huang and Thomas F. George

Department of Chemistry

University of Rochester

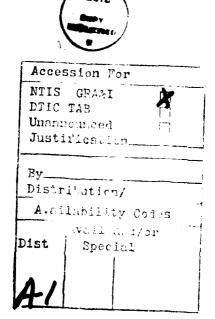
Rochester, New York 14627

Jian-Min Yuan and L. M. Narducci

Department of Physics and Atmosphere Science

Drexel University

Philadelphia, Pennsylvania 19104



<u>Abstract</u>

Vibrational excitation and relaxation for a molecule adsorbed on a surface is investigated in connection with desorption/dissociation processes. The population equations for a laser-driven anharmonic oscillator involving both energy- and phase-relaxation mechanisms are derived with the Zwanzig projector technique and eigenfunction-expansion method due to Weidlich. Phase relaxation is seen to assist in the excitation of the resonant active mode.

I. Introduction

Infrared laser radiation impinging onto a molecule adsorbed on a solid surface can resonantly excite the admolecular internal vibrational/rotational modes, or deposit photon energy in the adbond with which the molecule is bound to the surface. $^{1-5}$ It has been shown experimentally 6,7 that SF₆ and CH₂F molecules adsorbed on NaCl surfaces at low temperatures can be desorbed by the resonant excitation of the adsorbate internal vibration, namely, the v_2 mode. The treatment of the problem of many degrees of freedom for the molecule-surface system is very complicated and generally an insoluble problem. Interactions between the resonant active vibrational mode and the other inactive modes, and between the admolecule and surface excitations - surface phonons, surface plasmons, reflected photons (i.e., photons emitted by the admolecule and reflected by the surface), etc. - create a large number of dissipative channels by which absorption, dissociation, desorption or chemical reaction on the surface are strongly influenced. (Desorption is the breaking of the adbond between the admolecule and the surface, and dissociation is the breaking of a bond within the admolecule.) In this Letter, we are especially interested in the influence of surface-dephasing relaxation on the excitation of the resonant mode, which plays an important role in the process of desorption/dissociation.

II. Theory

Under normal circumstances, due to the vast number of degrees of freedom of the admolecule-surface system, the intrasystem relaxation can be very fast, and in fact, comparable to or faster than the rate of photon absorption. In this case, energy randomization or statistical thermodynamical behavior may at least partly occur in desorption/dissociation processes on surfaces. 8 The IR-laser-driven

resonant active mode absorbs photons and transfers its energy to other vibrational modes or the surface, and then "interference" superposition of these many vibrations can instantaneously accumulate energy in the weakest bond to lead to molecule dissociation or desorption from the surface, if this energy is beyond the appropriate the shold. 9,10 As an example, the threshold temperature for thermal dissociation of SF6 in the gas phase is measured to be 1600 K. 8 At this temperature, the average occupation number of the resonant mode (v_3) of SF6 is \overline{n} =0.76, which means that, while the energy deposited in the molecule has gone into the v_3 mode to exceed the critical value for dissociation, the other modes remain in low excited states.

We shall confine ourselves here to the excitation and relaxation of the active anharmonic vibrational mode which is driven by a resonant coherent field and controlled by both energy (T_1) and phase (T_2) relaxation. For this purpose, we treat the many other non-active modes and surface excitations (i.e., phonons, plasmons, reflected photons, etc.) as a large thermal reservoir, by which dissipation mechanisms are provided. A reduced density operator W for the resonant mode obeys the master equation

$$\frac{dW}{dt} = \left[\frac{dW}{dt}\right] \text{coherent} + \left[\frac{dW}{dt}\right] \text{incoherent}$$

$$= \left[-iL_c + \Lambda_1 + \Lambda_2\right] W. \tag{1}$$

The reversible term is $-iL_cW \equiv -(i/H)[H,W]$, where $H = H_M + H_F$ is the sum of anharmonic oscillator and external field interactions, 11

$$H_{M} = H(\omega - \omega_{L})a^{\dagger}a - H\epsilon a^{\dagger}a(a^{\dagger}a+1)$$
 (2)

$$H_{F} = \mu E(t)(a^{\dagger} + a). \tag{3}$$

Here a and a are oscillator ladder operators, μ is the molecular dipole moment, E(t) is the applied electric field, ω_L is the laser frequency, and ε is a positive anharmonicity. The dissipative terms are given by 12,13

$$\Lambda_1 W = \gamma_+([a,Wa^{\dagger}] + [aW,a^{\dagger}])$$

+ $\gamma_-([a^{\dagger},Wa] + [a^{\dagger}W,a]),$ (4)

$$\Lambda_2 W = n([a^{\dagger}a, Wa^{\dagger}a] + [a^{\dagger}aW, a^{\dagger}a]). \tag{5}$$

 Λ_2 W is responsible for a pure dephasing in the decay of the off-diagonal matrix elements, ¹⁴ and γ_+ , γ_+ and γ_+ contain implicit information about the reservoir, including the surface. With the dissipative terms [Eqs. (4) and (5) only], the relaxation behavior of the expectation values of the amplitude and number operator is governed by the equations $\frac{d}{dt}\langle a\rangle = -(\gamma_+ - \gamma_+ + \gamma_+)\langle a\rangle \text{ and } \frac{d}{dt}\langle a^+ a\rangle = 2(\gamma_+ - \gamma_+)\langle a^+ a\rangle + 2\gamma_+.$ Λ_2 thus affects only the off-diagonal elements of the density operator by providing a pure phase-destroying term with a rate of γ_+ . In the Born approximation, the diagonal parts of the density matrix obey the equation

$$\frac{d}{dt} W_{n} = \langle n | \Lambda_{1} W_{r} | n \rangle - \int_{0}^{t} d\tau \langle n | PL_{i} e^{-iL_{M}(t-\tau)} \Lambda_{2}^{(t-\tau)} L_{i} W_{r}(\tau) | n \rangle, \qquad (6)$$

where L_i is a Liouville operator due to the oscillator-field interaction, P is a projection operator used to obtain the diagonal part of the density operator, ¹⁵ and $W_r \equiv PW$ is the relevant part of the density matrix.

The difficult problem here is the explicit evaluation of the matrix elements that appear inside the integral in Eq. (6). We use the Weidlich eigenfuntion-expansion method 12,13 to evaluate these elements. The "eigenstate" of the super-operators Λ_2 and Λ_2^+ are $||A_{pq}||$ and $||B_{pq}||$, respectively, and under the Markoff assumption, the "eigenvalues" of Λ_2 have the form $\lambda_{pq} \approx -nq^2$, where p = 0, 1, 2, ... and q = 0, ±1, ±2, The equations of motion for the diagonal parts of the density matrix are then given as

$$\dot{W}_{n} = \langle n | \Delta_{1} W_{r} | n \rangle - \int_{0}^{t} d\tau \sum_{pq} e^{-nq^{2}(t-\tau)}$$

$$\times \text{Tr}(B_{pq}^{\perp} L_{1} W_{r}(\tau)) \langle n | L_{1} e^{-iL_{M}(t-\tau)} A_{pq} | n \rangle.$$
(7)

By a lengthy but straightforward algebraic calculation, we obtain a simpler set of coupled differential equations to describe the excitation and relaxation behavior for the resonant vibrational mode:

$$\dot{W}_{n} = 2\kappa \left[(n+1) W_{n+1} - nW_{n} \right] - 2\eta \left[\frac{\Omega_{R}(t)}{\eta} \right]^{2} \left\{ \frac{n(W_{n} - W_{n-1})}{1 + \left[2\frac{\varepsilon}{\eta}(n-1) - \frac{\Delta}{\eta} \right]^{2}} + \frac{(n+1)(W_{n} - W_{n+1})}{1 + \left(2\frac{\varepsilon}{\eta} n - \frac{\Delta}{\eta} \right)^{2}} \right\}.$$
(8)

The decay parameters κ ($\equiv \gamma_{\downarrow} - \gamma_{\uparrow}$) and n are determined by the internal mode-mode coupling and the admolecule-surface interaction, and for a fixed temperature and admolecule-surface distance they can be considered as constants (here a cold surface of 0 K has been assumed), $\Omega_{R}(t)$ is the Rabi frequency, and Δ is the detuning between the admolecular fundamental and laser frequencies.

The first two terms with the factor of 2κ on the right-hand side of Eq. (8) describe the energy relaxation processes which always lead to population loss in the resonant mode. In the next terms with the factors of n describing the influence of the phase relaxation, we notice that there are competitive factors: upon an increase of the value of n, the phase relaxation tends to diminish the detuning and anharmonicity "bottleneck" effects, and the effective Rabi frequency $\Omega_{\rm R}/n^{\frac{1}{2}}$ tends to decrease.

III. Calculations and Results

Figure 1 shows the influences of the phase relaxation. The time evolution for the average occupation number \overline{n} of a driven damped anharmonic oscillator is displayed, with consecutive values of the phase-relaxation parameter η (η = 1,5,9,13). We note that there is a range of values of η which assist the excitation of the admolecular active mode. For the optimal parameter η = 5 in our specific case, Figure 2 displays the time-dependent population distributions for several excited states, where considerable population is seen. One would generally expect the phase decay η to be much more enhanced than κ by the inactive adspecies vibrational modes and by the surface. The phase-dissipation mechanism may be generated in several ways in an admolecule-surface system as follows: (a) intramolecular coupling between the active and inactive vibrational modes; 17 (b) intermolecular collisions when the admolecules can migrate on the surface; 18 (c) rotational relaxation within the vibrational level manifold of the admolecule; 11 (d) self-coupling of the admolecule due to surface-reflected photons. 19,20

<u>Acknowledgments</u>

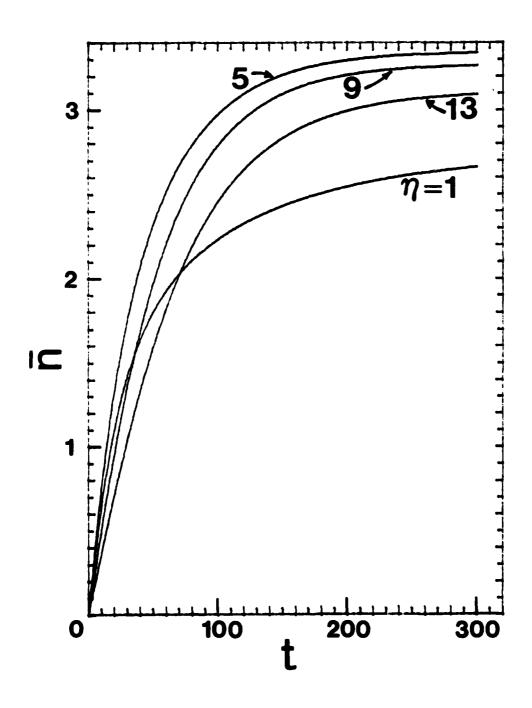
This research was supported in part by the Air Force Office of Scientific Research (AFSC), United States Air Force, under Grant AFOSR-82-0046, and the Office of Naval Research. The United States Government is authorized to reproduce and distribute reprints for governmental purposes notwithstanding any copyright notation hereon. TFG acknowledges the Camille and Henry Dreyfus Foundation for a Teacher-Scholar Award (1975-84).

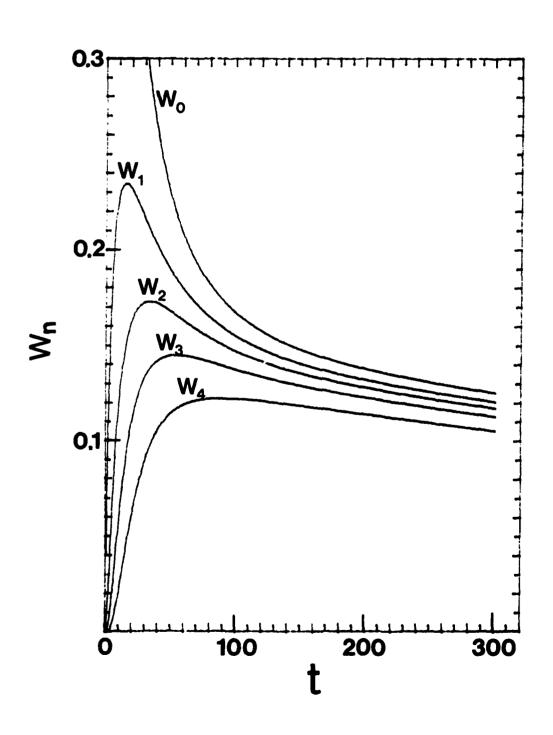
References

- 1. M. S. Slusky and T. F. George, Chem. Phys. Lett., <u>57</u> 474 (1978).
- 2. J. Lin and T. F. George, Surf. Sci., 100, 381 (1980).
- 3. J. Lin and T. F. George, J. Phys. Chem., 84, 2957 (1980).
- 4. D. Lucas and G. E. Ewing, Chem. Phys., <u>58</u>, 385 (1981).
- Z. W. Gortel, H. J. Kreuzer, P. Piercy and R. Teshima, Phys. Rev. B, <u>27</u>, 5066 (1983).
- J. Heidberg, H. Stein, A. Nestmann, E. Hoefs and I. Hussla, in <u>Laser-Solid</u>
 <u>Interactions and Laser Processing</u>, AIP Conf. Proc. 50, ed. by S. D. Ferris,
 H. J. Leamy and J. M. Poate (Am. Inst. Phys., New York, 1979), pp. 49-54.
- 7. J. Heidberg, H. Stein and E. Riehl, Phys. Rev. Lett. <u>49</u>, 666 (1982); ibid., Surf. Sci. 126, 183 (1983).
- For a similar case in the gas phase, see J. G. Black, E. Yablonovitch,
 N. Bloembergen and S. Mukamel, Phys. Rev. Lett., 38, 1131 (1977).
- 9. N. B. Slater, <u>Theory of Unimolecular Reactions</u> (Cornell University, Ithaca, 1959).
- 10. P. J. Robinson and K. A. Holbrook, <u>Unimolecular Reactions</u> (Wiley, New York, 1972), Chapt. 2.
- 11. L. M. Narducci, S. S. Mitra, R. A. Shatas and C. A. Coulter, Phys. Rev. A, <u>16</u>, 247 (1977).
- 12. W. Weidlich, Z. Phys., 241, 325 (1971).
- 13. X. Y. Huang, L. M. Narducci and J. M. Yuan, Phys. Rev. A 23, 3084 (1981).
- 14. A. H. Zewail, Acc. Chem. Res. <u>13</u>, 360 (1980).
- 15. R. Zwanzig, Physica (Ultrecht) 30, 1109 (1964).
- 16. P. L. Knight and P. W. Milonni, Physics Reports 66, 22 (1980).
- 17. J. Lin and T. F. George, Theor. Chim. Acta, <u>61</u>, 243 (1983).
- 18. X. Y. Huang, J. D. Cresser and J. H. Eberly, J. Chem. Phys., in press.
- 19. X. Y Huang, J. Lin and T. F. George, J. Chem. Phys., <u>80</u>, 893 (1984).
- 20. X. Y. Huang and T. F. George, J. Phys. Chem., in press.

Figure Captions

- 1. Time dependence of the average occupation number \overline{n} = $\langle a^{\dagger}a \rangle$ for (cw) laser-driven anharmonic oscillator in contact with a bath. The curves correspond to $\eta=1,\,5,\,9,\,13$, and the others parameters are $(\kappa,\,\Omega_R,\,\varepsilon,\,\Delta)$ = (0.01, 0.5, 1.5, 3.0), where all rates are in the unit of the transverse relaxation rate. This figure shows that when the vibrational linewidth η is smaller than the detuning Δ , the maximum vibrational excitation is small but grows as η increases. Excessively large values of η eventually reduce the average excitation level, presumably due to the spread of population within the linewidth.
- 2. The probability of occupation of level n, $W_n(t)$, as a function of time. η is fixed at 5, and the other parameters are given in Figure 1.





TECHNICAL REPORT DISTRIBUTION LIST, GEN

| | No. Copies | | No. Copies |
|--|---------------|--|---------------|
| Office of Naval Research Attn: Code 413 800 N. Quincy Street Arlington, Virginia 22217 | 2 | Naval Ocean Systems Center Attn: Technical Library San Diego, California 92152 | 1 |
| ONR Pasadena Detachment Attn: Dr. R. J. Marcus 1030 East Green Street Pasadena, California 91106 | | Naval Weapons Center Attn: Dr. A. B. Amster Chemistry Division China Lake, California 93555 | 1 |
| Commander, Naval Air Systems Command Attn: Code 310C (H. Rosenwasser) Washington, D.C. 20360 | 1 | Scientific Advisor Commandant of the Marine Corps Code RD-1 Washington, D.C. 20380 | 1 |
| Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko Port Hueneme, California 93401 | 1 | Dean William Tolles Naval Postgraduate School Monterey, California 93940 | 1 |
| Superintendent Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375 | 1 | U.S. Army Research Office Attn: CRO-AA-IP P.O. Box 12211 Research Triangle Park, NC 2770 | 1 |
| Defense Technical Information Cente Building 5, Cameron Station Alexandria, Virginia 22314 | r 12 | Mr. Vincent Schaper DTNSRDC Code 2830 Annapolis, Maryland 21402 | 1 |
| DTNSRDC Attn: Dr. G. Bosmajian Applied Chemistry Division Annapolis, Maryland 21401 | 1 | Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 1911 | 1 |
| Naval Ocean Systems Center Attn: Dr. S. Yamamoto Marine Sciences Division San Diego, California 91232 | 1 | Mr. A. M. Anzalone Administrative Librarian PLASTEC/ARRADCOM Bldg 3401 Dover, New Jersey 07801 | 1 |
| Dr. David L. Nelson Chemistry Program Office of Naval Research 800 North Quincy Street Arlington, Virginia 22217 | 1 | | |

TECHNICAL REPORT DISTRIBUTION LIST, 056

Dr. G. A. Somorjai Department of Chemistry University of California Berkeley, California 94720

Dr. J. Murday Naval Research Laboratory Surface Chemistry Division (6170) 455 Overlook Avenue, S.W. Washington, D.C. 20375

Dr. J. B. Hudson Materials Division Rensselaer Polytechnic Institute Troy, New York 12181

Dr. Theodore E. Madey Surface Chemistry Section Department of Commerce National Bureau of Standards Washington, D.C. 20234

Dr. Chia-wei Woo Department of Physics Northwestern University Evanston, Illinois 60201

Dr. Robert M. Hexter Department of Chemistry University of Minnesota Minneapolis, Minnesota

Dr. J. E. Demuth
IBM Corporation
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Or. M. G. Lagally
Department of Metallurgical
and Mining Engineering
University of Wisconsin
Madison, Wisconsin 53706

Dr. Adolph B. Amster Chemistry Division Naval Weapons Center China Lake, California 93555 Dr. W. Kohn
Department of Physics
University of California, San Diego
La Jolla, California 92037

Dr. R. L. Park
Director, Center of Materials
Research
University of Maryland
College Park, Maryland 20742

Dr. W. T. Peria Electrical Engineering Department University of Minnesota Minneapolis, Minnesota 55455

Dr. Keith H. Johnson
Department of Metallurgy and
Materials Science
Massachusetts Institute of Technology
Cambridge, Massachusetts 02139

Dr. J. M. White Department of Chemistry University of Texas Austin, Texas 78712

Dr. R. P. Van Duyne Chemistry Department Northwestern University Evanston, Illinois 60201

Dr. S. Sibener Department of Chemistry James Franck Institute 5640 Ellis Avenue Chicago, Illinois 60637

Dr. Arold Green Quantum Surface Dynamics Branch Code 3817 Naval Weapons Center China Lake, California 93555

Dr. S. L. Bernasek Princeton University Department of Chemistry Princeton, New Jersey 08544

TECHNICAL REPORT DISTRIBUTION LIST, 056

the state of the s

Dr. F. Carter Code 6132 Naval Research Laboratory Washington, D.C. 20375

Dr. Richard Colton Code 6112 Naval Research Laboratory Washington, D.C. 20375

Dr. Dan Pierce National Bureau of Standards Optical Physics Division Washington, D.C. 20234

Professor R. Stanley Williams Department of Chemistry University of California Los Angeles, California 90024

Dr. R. P. Messmer
Materials Characterization Lab.
General Electric Company
Schenectady, New York 22217
12301

Dr. Robert Gomer Department of Chemistry James Franck Institute 5640 Ellis Avenue Chicago, Illinois 60637

Dr. Ronald Lee R301 Naval Surface Weapons Center White Oak Silver Spring, Maryland 20910

Dr. Paul Schoen Code 5570 Naval Research Laboratory Washington, D.C. 20375

Dr. John T. Yates Department of Chemistry University of Pittsburgh Pittsburgh, Pennsylvania 15260 Or. Richard Greene Code 5230 Naval Research Laboratory Washington, D.C. 20375

Dr. L. Kesmodel
Department of Physics
Indiana University
Bloomington, Indiana 47403

Dr. K. C. Janda California Institute of Technology Division of Chemistry and Chemical Engineering Pasadena, California 91125

Professor E. A. Irene Department of Chemistry University of North Carolina Chapel Hill, Northc Carolina 27514

Or. Adam Heller Bell Laboratories Murray Hill, New Jersey 07974

Dr. Martin Fleischmann Department of Chemistry Southampton University Southampton \$09 5NH Hampshire, England

Dr. John W. Wilkins Cornell University Laboratory of Atomic and Solid State Physics Ithaca, New York 14853

Dr. Richard Smardzewski Code 6130 Naval Research Laboratory Washington, D.C. 20375

TECHNICAL REPORT DISTRIBUTION LIST, 056

Dr. R. G. Wallis Department of Physics University of California Irvine, California 92664

Dr. D. Ramaker Chemistry Department George Washington University Washington, D.C. 20052

Dr. P. Harsma
Physics Department
University of California
Santa Barbara, California 93106

Or. J. C. Hemminger Chemistry Department University of California Irvine, California 92717

Professor T. F. George Chemistry Department University of Rochester Rochester, New York 14627

Dr. G. Rubloff
TBM
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Professor Horia Metiu Chemistry Department University of California Santa Barbara, California 93106

Captain Lee Myers AFOSR/NC Bollig AFB Washington, D.C. 20332

Professor Roald Hoffmann Department of Chemistry Cornell University Ithaca, New York 14853 Dr. R. W. Plummer
Department of Physics
University of Pennsylvania
Philadelphia, Pennsylvania 19104

Dr. E. Yeager Department of Chemistry Case Western Reserve University Cleveland, Ohio 41106

Professor D. Hercules University Pittsburgh Chemistry Department Pittsburgh, Pennsylvania 15260

Professor N. Winograd
Department of Chemistry
Pennsylvania State University
University Park, Pennsylvania 16802

Dr. G. D. Stein Mechanical Engineering Department Northwestern University Evanston, Illinois 60201

Professor A. Stecki Department of Electrical and Systems Engineering Rensselaer Polytechnic Institute Troy, NewYork 12181

Professor G. H. Morrison Department of Chemistry Cornell University Ithaca. New York 14853

Dr. David Squire Army Research Office P.O. Box 12211 Research Triangle Park, NC 27709

